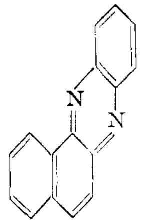
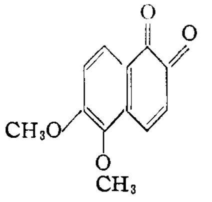
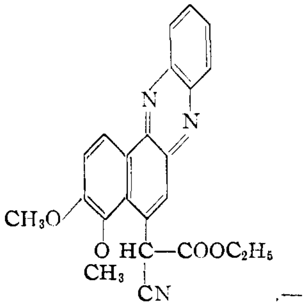
Marshall Gates. The Synthesis of Ring Systems Related to Morphine. III. 5,6-Dimethoxy-4-cyanomethyl-1,2-naphthoquinone and its Condensation with Dienes.

Pages 228 and 229. Formulas XI and XV should be


XI


XV

Also the formulas for the azine of compound XVI should have the structure


.—MARSHALL GATES

Hussein Sadek and Raymond M. Fuoss. Electrolyte-Solvent Interaction. I. Tetra-butylammonium Bromide in Methanol–Nitrobenzene Mixtures.

Page 304, Fig. 1. The methanol points should be at $-\log K = 1.42$ and $10^8a = 4.46$.

Page 305, Table IV. For No. 1, $K = 0.038$ and $a = 4.46$.—R. M. FUOSS.

John D. Roberts, Richard L. Webb and Elizabeth A. McElhill. The Electrical Effect of the Trifluoromethyl Group.

Page 409. In col. 1, text line 29, for "$C_8H_{10}NH_3$" read "$C_9H_{10}NF_3$." In text line 35, for "$C_8H_{10}NF_3$" read "$C_9H_{10}NF_3$."

Page 410. In Table IV, data line 2, for "*p*-CH_3" read "*p*-CF_3."—JOHN D. ROBERTS.

Martin J. Schick, Paul M. Doty and Bruno H. Zimm. Thermodynamics of Concentrated Polystyrene Solutions.

Page 531. In col. 2, line 13, the thermal expansion coefficient of cyclohexane should read "0.00123," instead of "0.0123."

Page 532. In col. 1 the number in line 7 of the text should be -1×10^{-2} mole-cm.^{9}/g.^{4}. In Table III the heading of the third column should be "$B_2 \times 10^4$ mole-cm.^{3}/g.^{2}" and the heating of the fourth column should be "$B_3 \times 10^3$ mole-cm.^{6}/g.^{3}."

The values given of the virial coefficients represent our best estimates *assuming that the higher virial coefficients are not zero*. If we had attempted to fit the data in the best manner with only three or four coefficients, somewhat different values would have been found.

We regret having omitted a reference to the work of George V. Browning and John D. Ferry, *J. Chem. Phys.*, 17, 1107 (1949).—PAUL DOTY and BRUNO ZIMM.

C. F. H. Allen, Jean V. Crawford, R. H. Sprague, Eleanor R. Webster and C. V. Wilson. 3-Azabenzanthrone Dyes.

Page 586. In Table I, data line 8, for "$C_{27}H_{10}F_3N_2O_4$" read "$C_{27}H_{19}F_3N_2O_4$."—C. F. H. ALLEN.

Richard C. Gilmore, Jr., and W. J. Horton. The Synthesis of 1-Hydroxycyclohepta[de]naphthalene.

Pages 733 ff. The authors write: "In every case where -cyclohepta[de]naphthalene is mentioned, it should be -7,8,9,10-tetrahydrocyclohepta[de] naphthalene, to conform to *Chemical Abstracts* nomenclature."—W. J. HORTON.

K. J. Palmer and Merle Ballantyne. The Structure of (I) Some Pectin Esters and (II) Guar Galactomannan.

Page 737. In col. 2, text line 1, for "Table II" read "Table I."

Robert F. Raffauf. Furans. II. Ultraviolet Absorption and Molecular Structure.

Page 754. In the legend of Fig. 1, for "*n*" read "*c*." Omit ⊖ on Formula I, change ⊖ to right side of Formula I, second resonance form, and insert ⊖ in place of — on Formula IV, second resonance form. In Formula VI, place the ⊕ on the α-carbon.

Page 755. In footnote (12) for "with" read "when."—ROBERT F. RAFFAUF.

William S. Johnson and Robert P. Graber. The Stobbe Condensation with 6-Methoxy-2-propionylnaphthalene. A Synthesis of Bisdehydrodoisynolic Acid.

Page 931. In col. 2, line 27, for "H" read "CH_3."—WILLIAM S. JOHNSON.

Charles A. Burkhard. The Reaction of Mercaptans with Alkenyl Silanes.

Page 1078. In Table I, Compound VIII, Formula "$(CH_3)_3SiOSi(CH_2)_3(CH_2)_3SCH_2CO_2H$," should read "$(CH_3)_3SiOSi(CH_3)_2(CH_2)_3SCH_2CO_2H$." Compound IX, Formula "$(C_2H_5O)_3Si(CH_3)_3SCH_2CO_2C_2H_5$," should read "$(C_2H_5O)_3Si(CH_2)_3SCH_2CO_2C_2H_5$."—CHARLES A. BURKHARD.

W. von E. Doering and Jerome A. Berson. A Reexamination of the Diisohomogenol Structure.

Page 1120. In col. 1, line 26, for "$C_{22}H_{26}O_6$" read "$C_{22}H_{26}O_7$."

Page 1121. In col. 2, line 17, for "C, 68.37" read "C, 68.87."—JEROME A. BERSON.

Marshall Gates, R. B. Woodward, W. F. Newhall and Rosemarie Künzli. The Synthesis of Ring Systems Related to Morphine. IV. N-Methylisomorphinane.

Page 1144. In column 2, the next to last line of the description of the preparation of **4-cyanomethyl-1,2-naphthoquinone (XV)**, (16%) should read (61%), and in the fifth line of the description of the preparation of **9,10-dioxo-13-cyanomethyl-5,8,9,10,13,14-hexahydrophenanthrene** 11.5 g. should read 15.9 g.—MARSHALL GATES.

R. A. Day, Jr., A. E. Robinson, Jr., J. M. Bellis and S. B. Till. Oxidation Potentials and Ultraviolet Absorption Spectra of a Series of Normal Ketones.

Page 1379. The authors write: "The data on the ultraviolet absorption spectra of the aliphatic ketones are not reliable, because the ketones were not spectroscopically pure. Rice (*Proc. Roy. Soc.* (*London*), **91A**, 76 (1914)), in an article which we carelessly overlooked, proved that when a saturated aliphatic ketone showed absorption on the ultraviolet side of the carbonyl band, a highly absorbing impurity, not removed by distillation, was present. Chemical purification seems to be necessary in order to remove this impurity. No quantitative significance should be given to our results. However, the qualitative conclusions which were made in our article seem to be valid. Rice reported that the wave length of maximum absorption shifted from 275 mμ for acetone to 282 mμ for di-*n*-propyl ketone. The only significant changes he found in molar absorption indices were between acetone (17.1) and methyl ethyl ketone (19.4). All larger ketones showed a constant value of 21.2."—R. A. DAY, JR.